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table using  $s p d f$  notation what is the electron configuration for a neutral

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$2s^2 2p^1$   $3s^2 3p^2$   $4s^2 4p^1$   $5s^2 5p^2$   $6s^2 6p^1$   $7s^2 7p^2$  show periodic table stuck

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configurations as shown

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number of 8 give its electron

configuration answer an atomic number

of 8 means a total of eight electrons

two of these electrons can fit into the

first shell while the remaining six can

fit into the second shell therefore the

configuration is written as  $2s^2 2p^6$

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$s p d$  and  $f$  have different shapes and

one orbital can hold a maximum of two electrons the  $p d$  and  $f$  orbitals have different sublevels thus can hold more electrons as stated the electron

configuration of each element is unique to its position on the periodic table

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fit into the  $1s$  orbital  $1s^2$  then next two

into the  $2s$  orbital  $2s^2$  the  $2p$  orbital

has three slots which can take 2

electrons each a total of 6 which gives

$2p^6$  the final two electrons fit in the  $3s$

orbital  $3s^2$  electron configuration

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web using the periodic table to

determine the electron configurations

of atoms is key but also keep in mind

that there are certain rules to follow

when assigning electrons to different

orbitals the periodic table is an

incredibly helpful tool in writing

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